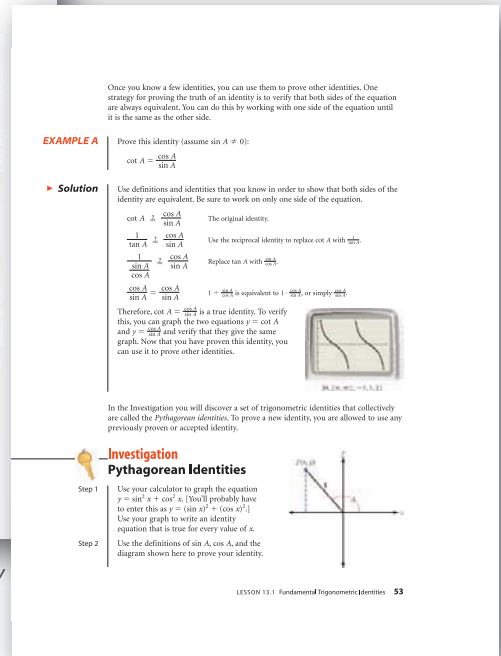
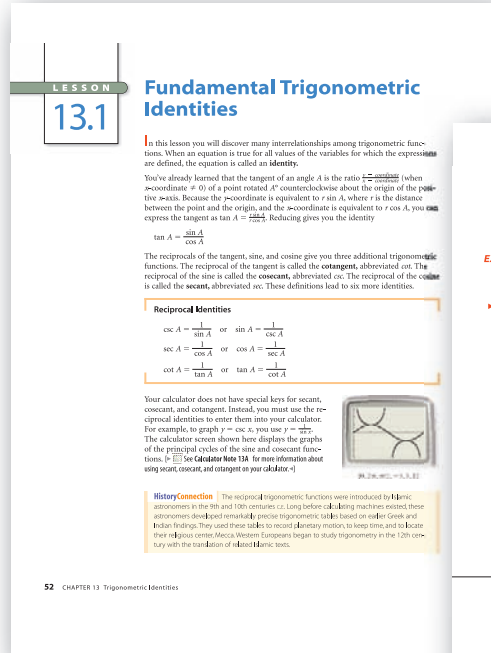


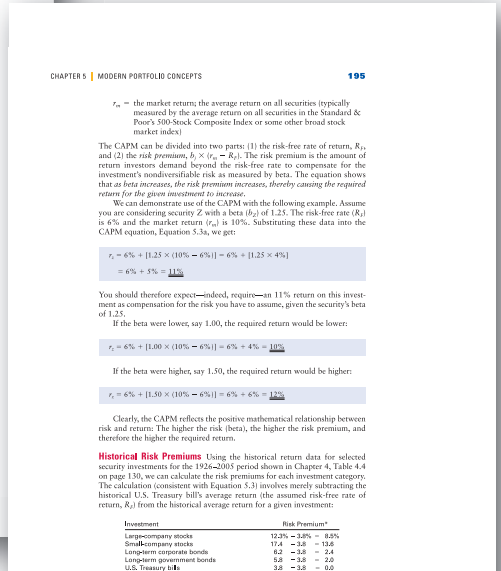
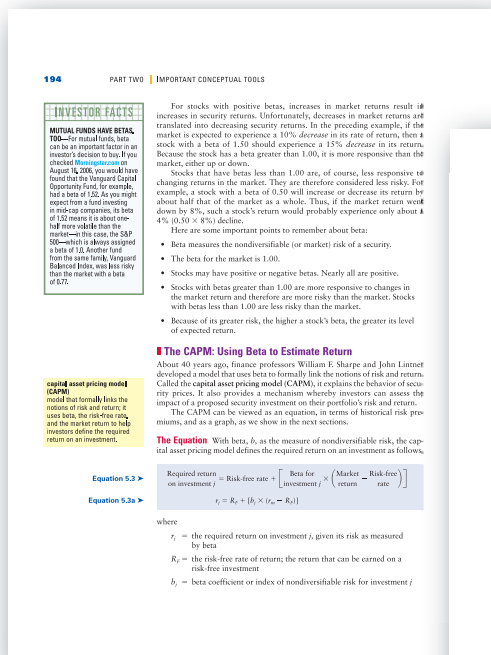
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Chapter 5 Reactions in Aqueous Solution

Finally, notice that there must be a charge balance as well as a mass balance in a balanced chemical equation. In the $\text{Ag}^+ + \text{Cl}^-$ net ionic equation, the **total** and **anion** charges on the left add together to give a net charge of 0, the same as the **net** charge on $\text{AgCl}(s)$ on the right.

Example 5.3—Writing and Balancing Net Ionic Equations

See the General ChemistryNow CD-ROM or website:

• Screen 5.7 Net Ionic Equations, for a tutorial on writing net ionic equations

Example 5.3—Writing and Balancing Net Ionic Equations

Problem Write a balanced, net ionic equation for the reaction of aqueous solutions of BaCl_2 and Na_2SO_4 to give BaSO_4 and NaCl .

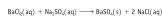
Strategy First, write a balanced equation for the overall reaction. Next, decide which compounds are soluble in water (Figure 5.3) and determine the ions that these compounds produce in solution. Finally, eliminate ions that appear on both the reactant and product sides of the equation.

Solution

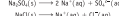
Step 1. Write the balanced equation.



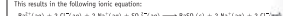
Step 2. Decide on the solubility of each compound. Compounds containing sodium ions are always water-soluble, and those containing chloride ions are almost always soluble. Soluble salts are also usually soluble, with one important exception being BaSO_4 . We can therefore write



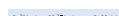
Step 3. Identify the ions in solution. All soluble ionic compounds dissociate to form ions in aqueous solution. (All are electrolytes.)



This results in the following ionic equation:



Step 4. Identify and eliminate the spectator ions (Na^+ and Cl^-) to give the net ionic equation.



Notice that the sum of ion charges is the same on both sides of the equation. On the left, $2+$ and $2-$ give zero; on the right, the charge on BaSO_4 is also zero.

Comment The steps followed in this example represent a general approach to writing net ionic equations.

Balancing Redox

When an ionic compound with halide ions dissolves in water, the halide ions are released into aqueous solution. Thus, BaCl_2 produces one Ba^{2+} ion and two Cl^{-} ions for each Ba^{2+} ion (and not Cl_2 or Cl_2^{2-} ions).



Precipitation reaction. The reaction of barium chloride and sodium sulfate produces insoluble barium sulfate and water-soluble sodium chloride. See Example 5.3.

Chemistry & Chemical Reactivity
John C. Kotz
Thomson Brooks/Cole 2006

Introduction to Econometrics
Stock/Watson
Addison Wesley (2006)

194 CHAPTER 6 Linear Regression with Multiple Regressors

Key Concept

6.2

THE MULTIPLE REGRESSION MODEL

The multiple regression model is

$$Y_i = \beta_0 + \beta_1 X_{i1} + \beta_2 X_{i2} + \cdots + \beta_k X_{ik} + u_i, i = 1, \dots, n \quad (6.7)$$

where

- Y_i is the i th observation on the dependent variable; $X_{i1}, X_{i2}, \dots, X_{ik}$ are the i th observations on each of the k regressors; and u_i is the error term.
- The population regression line is the relationship that holds between Y and the X 's on average in the population:

$$E(Y|X_1 = x_1, X_2 = x_2, \dots, X_k = x_k) = \beta_0 + \beta_1 x_1 + \beta_2 x_2 + \cdots + \beta_k x_k$$

- β_1 is the slope coefficient on X_1 ; β_2 is the coefficient on X_2 , and so on. The coefficient β_1 is the expected change in Y resulting from changing X_1 by one unit, holding constant X_2, \dots, X_k . The coefficients on the other X 's are interpreted similarly.
- The intercept β_0 is the expected value of Y when all the X 's equal 0. The intercept can be thought of as the coefficient on a regressor, X_0 , that equals 1 for all i .

of practical help to the superintendent, however, we need to provide her with estimates of the unknown population coefficients β_0, \dots, β_k of the population regression model calculated using a sample of data. Fortunately, these coefficients can be estimated using ordinary least squares.

6.3 The OLS Estimator in Multiple Regression

This section describes how the coefficients of the multiple regression model can be estimated using OLS.

$$\text{Required return on investment } j = \text{Risk-free rate} + \left[\text{Beta for investment } j \times \left(\text{Market return} - \text{Risk-free rate} \right) \right]$$

$$r_j = R_F + [b_j \times (r_m - R_F)]$$

+ $b_k X_{ki}$, and the mistake in predicting Y_i is $Y_i - (b_0 + b_1 X_{1i} + \cdots + b_k X_{ki}) = Y_i - b_0 - b_1 X_{1i} - \cdots - b_k X_{ki}$. The sum of these squared prediction mistakes over all n observations thus is

$$\sum_{i=1}^n (Y_i - b_0 - b_1 X_{1i} - \cdots - b_k X_{ki})^2. \quad (6.8)$$

$$\frac{\cos A}{\sin A} = \frac{\cos A}{\sin A}$$

$$1 \div \frac{\sin A}{\cos A} \text{ is equivalent to } 1 \cdot \frac{\cos A}{\sin A}, \text{ or simply } \frac{\cos A}{\sin A}.$$

Therefore, $\cot A = \frac{\cos A}{\sin A}$ is a true identity. To verify this, you can graph the two equations $y = \cot A$ and $y = \frac{\cos A}{\sin A}$ and verify that they give the same



Step 3. Identify the ions in solution. All soluble ionic compounds dissociate to form ions in aqueous solution. (All are electrolytes.)

